

Data for Experiment 4

Record your observations for each combination below. If a reaction occurs, write balanced MOLECULAR, IONIC, and NET-IONIC equations. If no reaction occurs, write NR. Make sure to include the physical states of all the products.

1. NaCl(aq) and $\text{KNO}_3\text{(aq)}$

Observations:

Molecular:

Ionic:

Net-Ionic:

2. NaCl(aq) and $\text{AgNO}_3\text{(aq)}$

Observations:

Molecular:

Ionic:

Net-Ionic:

3. $\text{Na}_2\text{CO}_3\text{(aq)}$ and HCl(aq)

Observations:

Molecular:

Product of *secondary* reaction:

4. NaOH(aq) and HCl(aq)

Observations:

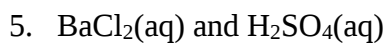
Molecular:

Ionic:

Net-Ionic:

Name: _____

Section: _____

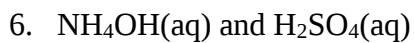


Observations:

Molecular:

Ionic:

Net-Ionic:

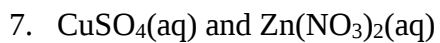


Observations:

Molecular:

Ionic:

Net-Ionic:

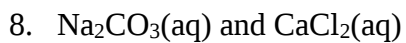


Observations:

Molecular:

Ionic:

Net-Ionic:



Observations:

Molecular:

Ionic:

Net-Ionic:

Name: _____

Section: _____

9. $\text{CuSO}_4(\text{aq})$ and $\text{NH}_4\text{Cl}(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

10. $\text{NaOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

11. $\text{FeCl}_3(\text{aq})$ and $\text{NH}_4\text{OH}(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic: