

Data for Experiment 7

Record your observations for each combination below. If a reaction occurs, write balanced MOLECULAR, IONIC, and NET-IONIC equations. If no reaction occurs, write NR. Make sure to include the physical states of all the products.

1. NaCl(aq) and $\text{KNO}_3\text{(aq)}$

Observations:

Molecular:

Ionic:

Net-Ionic:

2. NaCl(aq) and $\text{AgNO}_3\text{(aq)}$

Observations:

Molecular:

Ionic:

Net-Ionic:

3. $\text{Na}_2\text{CO}_3\text{(aq)}$ and HCl(aq)

Observations:

Molecular:

Ionic:

Net-Ionic:

4. NaOH(aq) and HCl(aq)

Observations:

Molecular:

Ionic:

Net-Ionic:

Name: _____

Section: _____

5. $\text{BaCl}_2(\text{aq})$ and $\text{H}_2\text{SO}_4(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

6. $\text{NH}_4\text{OH}(\text{aq})$ and $\text{H}_2\text{SO}_4(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

7. $\text{CuSO}_4(\text{aq})$ and $\text{Zn}(\text{NO}_3)_2(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

8. $\text{Na}_2\text{CO}_3(\text{aq})$ and $\text{CaCl}_2(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

Name: _____

Section: _____

9. $\text{CuSO}_4(\text{aq})$ and $\text{NH}_4\text{Cl}(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

10. $\text{NaOH}(\text{aq})$ and $\text{HNO}_3(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

11. $\text{FeCl}_3(\text{aq})$ and $\text{NH}_4\text{OH}(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

12. $\text{Na}_2\text{SO}_3(\text{aq})$ and $\text{HCl}(\text{aq})$

Observations:

Molecular:

Ionic:

Net-Ionic:

Questions

1. For each of the reactions listed below, write balanced molecular, ionic, and net-ionic equations. If no reaction occurs, write NR. Assume all reactants are aqueous unless otherwise noted. Include all physical states.

A. Lead(II) nitrate and magnesium sulfate solutions are combined.

Molecular:

Ionic:

Net-Ionic:

B. Barium chloride solution is poured into a solution of ammonium carbonate.

Molecular:

Ionic:

Net-Ionic:

C. Magnesium chloride solution is mixed with nickel(II) nitrate solution.

Molecular:

Ionic:

Net-Ionic:

D. Cobalt(II) sulfate and lithium sulfide solutions are combined.

Molecular:

Ionic:

Net-Ionic:

E. Hydrochloric acid solution is reacted with a solution of lithium carbonate.

Molecular:

Ionic:

Net-Ionic:

Name: _____

Section: _____

F. Hydroiodic acid and ammonium sulfite solutions are mixed.

Molecular:

Ionic:

Net-Ionic:

G. Sodium hydroxide solution is poured into a solution of cobalt(II) chloride.

Molecular:

Ionic:

Net-Ionic:

H. Ammonium chloride and potassium hydroxide solutions are reacted.

Molecular:

Ionic:

Net-Ionic:

I. Solid strontium bromide is mixed with a solution of potassium phosphate.

Molecular:

Ionic:

Net-Ionic:

J. Solutions of ammonium sulfate and sodium chloride are combined.

Molecular:

Ionic:

Net-Ionic: